



Dr. Atkins'  
**Diet Revolution**  
**THE HIGH CALORIE WAY**  
**TO STAY THIN FOREVER**  
Robert C. Atkins, M.D.  
**THE FASHION VOGUE SUPERDIET**  
**EXPLAINED IN FULL.**

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# Diet Revolution

**RM Cervero**



## **Diet Revolution:**

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web 1 consider the reaction  $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$  a 80.0 grams of iodine v oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced b if in the above situation only 0.160 moles of iodine  $\text{I}_2$  was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the

theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$
  
 8.3 mole  $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield =  $137.1 \text{ g } \text{H}_2\text{O} / 149.4 \text{ g } \text{H}_2\text{O} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
 if the reaction of 21.3 grams of  $\text{O}_2$  produces a 4.73 g yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  
 $21.3 \text{ g } \text{O}_2 / 32 \text{ g/mole } \text{O}_2 = 0.666 \text{ moles } \text{O}_2$   
 $0.666 \text{ moles } \text{O}_2 \times 3 \text{ g CO} / \text{mole } \text{O}_2 = 1.998 \text{ g CO}$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  

$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$
  
 b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 / \text{mol Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$   
 c determine the percent yield  
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$   
 theoretical = 302.3 g

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web 7. Na<sub>2</sub>CO<sub>3</sub> is the limiting reactant percent yield = 86.68% only four molecules can be made 9. this amount cannot be weighed by ordinary balances and is worthless 10. nitrogen is the limiting reagent 11. yes methane is the limiting reagent 12.

c is the limiting reagent 4 33 g of h<sub>2</sub> are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3