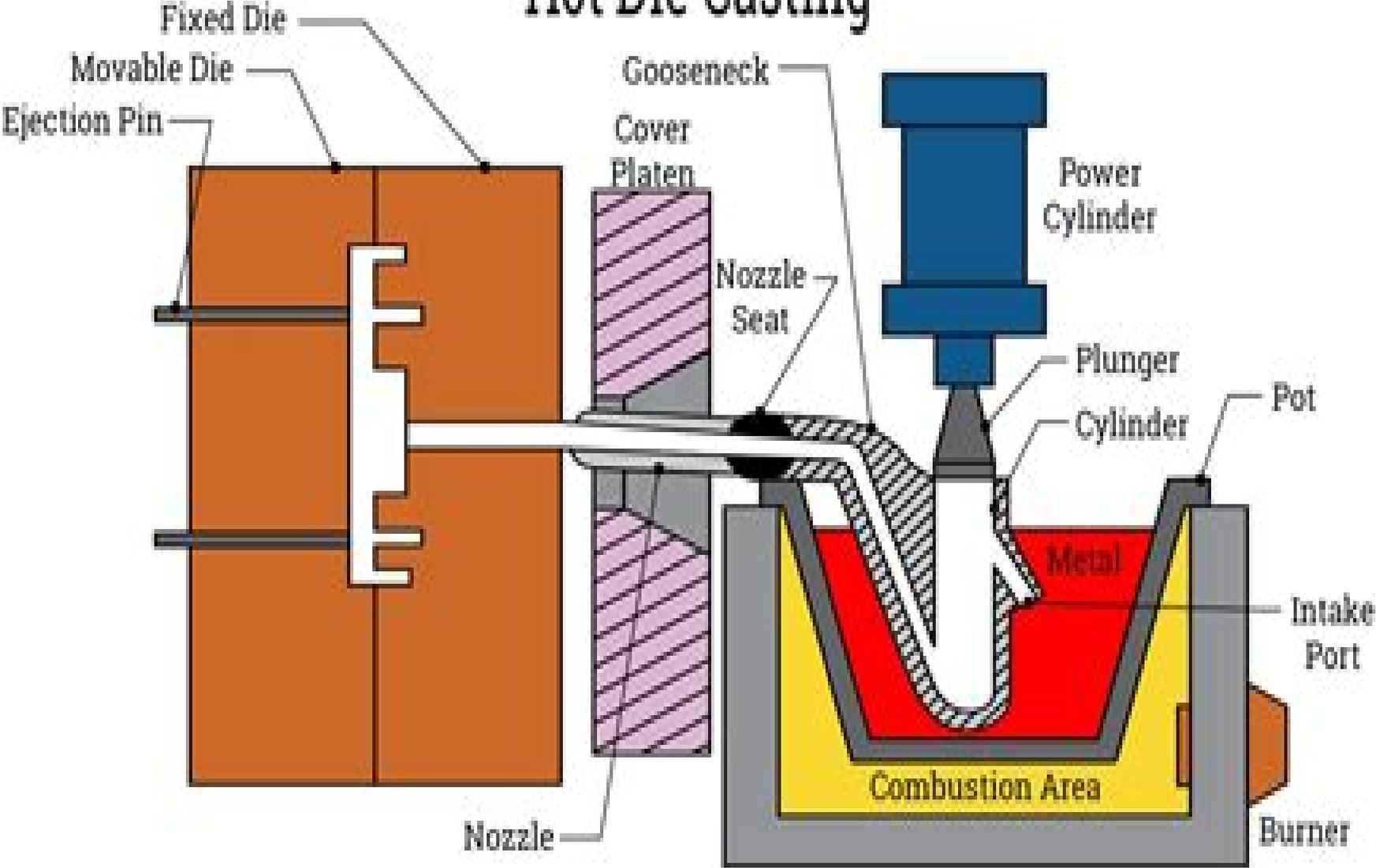


# Hot Die Casting



# Die Casting Engineering

**Jiyuan Zhang**



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subsequent  $\Delta h$  values  $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} + \text{C}_2\text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_4 \text{ g} + \text{O}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g}$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} \rightarrow \frac{1}{2} \text{O}_2 \text{ g} + \text{H}_2\text{O l}$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g}$

co<sub>2</sub> g 3 h<sub>2</sub> o g Δh 283 kj g 1 c<sub>2</sub>h<sub>6</sub> g

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web thus for reaction 9 2 in which 1 mole of  $h_3o^+$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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