

**PROBLEM SET 0: REVIEW**  
**CHEM 101 Foundations Sept. 3, 2025**

**References:** Chapters 1–4 Definitions, elements and compounds, stoichiometry, classes of reactions

- (OpenStax, Problem 1.27) Classify as physical or chemical changes:  
(a) Condensation of steam  
(b) Burning of gasoline  
(c) Sealing of milk  
(d) Dissolving of sugar in water  
(e) Melting of gold
- (OpenStax, Problem 1.53) Perform these calculations and report to the correct number of significant figures.  
(a)  $628 \times 342$   
(b)  $(5.63 \times 10^2) \times (7.4 \times 10^3)$   
(c)  $28.0 / 13.483$   
(d)  $8119 \times 0.000023$   
(e)  $14.98 + 27.340 + 84.7593$   
(f)  $42.7 + 0.259$
- (Silberberg, Problem 2.58) Refer to a periodic table.  
(a) The heaviest alkaline-earth metal has the symbol \_\_\_ and atomic number \_\_\_  
(b) The lightest group 14 metalloid has the symbol \_\_\_ and atomic number \_\_\_  
(c) The coinage metal (group 11) whose atoms have the fewest electrons has symbol \_\_\_ and atomic mass \_\_\_  
(d) The halogen in period 4 has the symbol \_\_\_ and atomic mass \_\_\_
- (Silberberg, Problem 2.132) Antimony is used in infrared devices and lead storage batteries. The element occurs in two naturally occurring isotopes, one with mass 120.904 amu and the other with mass 122.904 amu.  
(a) Write the  ${}^A_ZX$  notation for each isotope, giving the mass number and numbers of  $n^0$ ,  $p^+$ , and  $e^-$ .  
(b) Use the atomic mass of antimony from the periodic table to calculate the natural abundance of each isotope.
- Name the compound or write the formula.  
(a)  $\text{HClO}_4$   
(b)  $\text{KNO}_3$   
(c)  $\text{Na}_2\text{Cr}_2\text{O}_7$   
(d)  $\text{S}_2\text{N}_4$   
(e) an oxyacid of nitrogen with N in the oxidation state of +3  
(f) lead(II) acetate  
(g) dinitrogen tetroxide  
(h) sodium hydrogen sulfite  
(i) phosphorous acid  
(j) a chlorine-containing anion in which the oxidation state of Cl is +3 and the ionic charge is 3-
- (OpenStax, Problem 4.84) The principal component of mothballs is naphthalene, a compound containing carbon and hydrogen with a molecular mass of about 130 g/mol. A 3.000-mg sample of naphthalene burns to give 10.3 mg of  $\text{CO}_2$ . Determine its empirical and molecular formulas.
- (OpenStax, Problem 4.73) Uranium can be isolated from its ores by dissolving it as  $\text{UO}_2(\text{NO}_3)_2$ , then separating it as solid  $\text{UO}_2(\text{C}_2\text{O}_4) \cdot 3\text{H}_2\text{O}$ . Addition of 0.4031 g of sodium oxalate,  $\text{Na}_2\text{C}_2\text{O}_4$ , to a solution containing 1.481 g of uranyl nitrate,  $\text{UO}_2(\text{NO}_3)_2$ , yields 1.073 g of solid  $\text{UO}_2(\text{C}_2\text{O}_4) \cdot 3\text{H}_2\text{O}$  according to the unbalanced equation:  
$$\text{Na}_2\text{C}_2\text{O}_4 + \text{UO}_2(\text{NO}_3)_2 + \text{H}_2\text{O} \rightarrow \text{UO}_2(\text{C}_2\text{O}_4) \cdot 3\text{H}_2\text{O} + \text{NaNO}_3$$
  
Balance the equation, and determine the limiting reactant and the percent yield of this reaction.

# Chemskill Foundations Personalized Problem Sets For Introductory Chemistry

**Minjie Lin**



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