

Chemistry Study Guide

Periodic Table of the Elements

| | | | | | | | | | | | | | | | | | | | | | | | | | | | |
|----|----|----|----|----|----|----|----|----|----|----|----|-----|----|-----|----|-----|-----|---|----|----|----|----|----|----|----|----|----|
| 1 | 2 | | | | | | | | | | | 3 | 4 | 5 | 6 | 7 | 8 | 9 | 10 | 11 | 12 | 13 | 14 | 15 | 16 | 17 | 18 |
| H | He | | | | | | | | | | | B | C | N | O | F | Ne | | | | | | | | | | |
| Li | Be | | | | | | | | | | | Al | Si | P | S | Cl | Ar | | | | | | | | | | |
| Na | Mg | Sc | Ti | V | Cr | Mn | Fe | Co | Ni | Cu | Zn | Ga | Ge | As | Se | Br | Kr | | | | | | | | | | |
| Rb | Sr | Y | Zr | Nb | Mo | Tc | Ru | Rh | Pd | Ag | Cd | In | Sn | Sb | Te | I | Xe | | | | | | | | | | |
| Cs | Ba | | Hf | Ta | W | Re | Os | Ir | Pt | Au | Hg | Tl | Pb | Bi | Po | At | Rn | | | | | | | | | | |
| Fr | Ra | | Rf | Db | Sg | Bh | Hs | Mt | Ds | Rg | Cn | Uut | Fl | Uup | Lv | Uus | Uuo | | | | | | | | | | |
| | | La | Ce | Pr | Nd | Pm | Sm | Eu | Gd | Tb | Dy | Ho | Er | Tm | Yb | Lu | | | | | | | | | | | |
| | | Ac | Th | Pa | U | Np | Pu | Am | Cm | Bk | Cf | Es | Fm | Md | No | Lr | | | | | | | | | | | |

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

Matter:

Anything that has **mass** and **occupies space** 🌟🌍

Examples: air, gases, liquids, solidz!

Atoms:

Atoms are the **smallest, indivisible particle** 🔬👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

| Particle | Symbol | Mass (amu) | Charge | Location |
|----------|--------|------------|--------|-------------|
| Proton | p^+ | 1 | +1 | Nucleus |
| Neutron | n | 1 | 0 | Nucleus |
| Electron | e^- | ~ 0 | -1 | Orbit/Shell |

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