



PATRICK E. ANDREWS

DESPERADO RUN

HE WAS THE LAST OF THE OLD GUARD
OUTLAWS — WITH ONE LAST
CHANCE TO LIVE!

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Desperado Run

O García



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determine based on electronegativity and symmetry if the following molecules are polar or nonpolar HCN $\text{H}_2\text{C}_2\text{N}_2$ CF_4 COCl_2 CO_2 NF_3 N_2 O_3

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web the relative electronegativities of the bonded atoms is $\text{H} < \text{C} < \text{Cl}$ and so the bond moments all point toward the Cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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covalent electronegativity difference is less than 0.4 nonmetal nonmetal close together on the periodic table

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web apr 1 2021 updated on april 01 2021 chemical bonds may be classified as being either polar or nonpolar the difference is how the electrons in the bond are arranged key takeaways what is a polar bond in chemistry a polar bond is a type of covalent bond in which the electrons forming the bond are unequally distributed