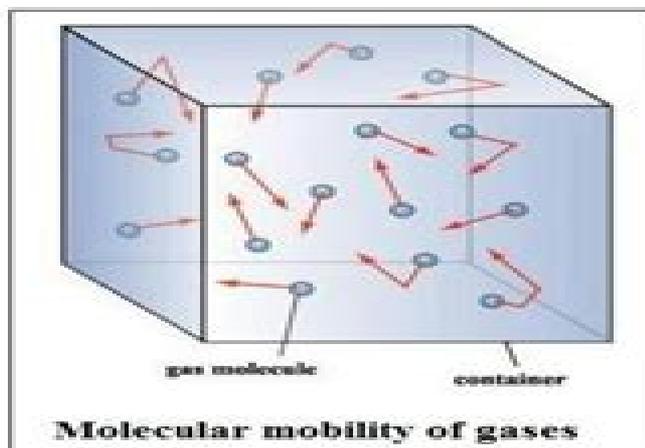


BASIC CHEMICAL PRINCIPLES

Behavior of ideal gases:

Kinetic theory of gases:

A gas is believed to be composed of molecules. These molecules are considered to be material particles and are separated by a large distance when compared to that of the molecules of solid and liquids. According to "Newton's law of motion of material bodies", the gas molecules are free to move in space. It is also assumed that the gas molecules behave like a "perfect elastic sphere". They neither gain nor lose energy during collisions. The size of the gas molecule is small when compared to the distance between the molecules and the volume occupied by a gas is large when compared to the volume of the individual gas molecule.



Kinetic Theory of Gases Postulates:

The kinetic theory of gas postulates is useful in the understanding of the macroscopic properties from the microscopic properties.

- Gases consist of a large number of tiny particles (atoms and molecules). These particles are extremely small compared to the distance between the particles.
- The size of the individual particle is considered negligible and most of the volume occupied by the gas is empty space. This postulate leads to the fact that a gas can be compressed.
- These molecules are in constant random motion which results in colliding with each other and with the walls of the container. These collisions with the walls cause the pressure exerted by the gas.

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