

STUDY CARD for CAMPBELL BIOLOGY, Ninth Edition

Reece • Urry • Cain • Wasserman • Minorsky • Jackson

Themes in the Study of Life (Ch. 1)

1. The theme accounts for life's unity and diversity.
2. New perspectives emerge at each level of the biological hierarchy, from atoms to the biosphere.
3. Organisms interact with their environments.
4. Life requires energy transfer.
5. Structure and function are correlated.
6. The cell is the basic unit of an organism.
7. DNA encodes heritable information.
8. Feedback mechanisms regulate biological systems.

The Chemical Context of Life (Ch. 2)

1. Matter consists of chemical elements or pure forms and is chemically defined compounds.
2. An element's properties depend upon the structure of its atoms.
 - a. The number of protons (Z) defines the element.
 - b. Electrons (Z) occupy electron shells.
3. Chemical bonds form when atoms interact and complete their outer electron shells.
 - a. Single or double covalent bonds form when pairs of electrons are shared between atoms.
 - b. Electrons transfer between atoms. An ionic bond is the attraction between oppositely charged ions.
 - c. Weak bonds include hydrogen bonds and van der Waals forces.
4. Chemical reactions occur and bond chemical bonds or reactants are converted to products.

Water and Life (Ch. 3)

1. Water molecules are polar. Hydrogen bonding between water molecules gives water unique properties that allow life to exist on Earth.
 - a. Cohesion due to hydrogen bonding contributes to the transport of water against gravity in plants.
 - b. Water's high specific heat and high heat of vaporization moderate temperatures on Earth.
 - c. Ice floats because hydrogen bonds lock water into a crystalline lattice that reduces its density.



- a. Water is a versatile solvent.
 - b. Hydrophilic substances have an affinity for water whereas hydrophobic substances repel water.
2. Acids and bases contribute to pH regulation.
 - a. Acidic solutions have a pH < 7 .
 - b. Basic solutions have a pH > 7 .
 - c. Buffers minimize changes in pH.

Carbon and Molecular Diversity (Ch. 4)

1. Organic chemistry is the study of carbon compounds. Carbon atoms can form diverse molecules by bonding with up to four other atoms.
 - a. Carbon chains form the skeletons of most organic molecules.
 - b. Hydrocarbons consist of only C and H.
 - c. Isomers have the same molecular formula but different structures and properties.
2. Functional groups are the parts of molecules involved in chemical reactions.

3. ATP is an important source of energy for cellular processes: $ATP \rightarrow ADP + P_i + \text{Energy}$.

Large Biological Molecules (Ch. 5)

1. Macromolecules are polymers synthesized from monomers by dehydration reactions.
 - a. Polymers are dissociated by monomers by the addition of a water molecule (hydrolysis).
 - b. Variation in the linear sequence of monomers results in the diversity of polymers.
 - c. Four major classes of large biological molecules are found in cells.
2. Carbohydrates include sugars, starch, and cellulose (and) and polymers of sugars (polysaccharides).
 - a. Most monosaccharides are multiples of CH_2O .
 - b. Starch (in plants) and glycogen (in animals) are polysaccharides that store energy.
 - c. Cellulose (in plants) and chitin (in animals) are polysaccharides that provide structural support.
3. Lipids are hydrophobic.
 - a. Triglycerides linked to three fatty acid chains can be saturated or unsaturated; they store energy.
 - b. Phospholipids (glycerol, two fatty acids, and phosphate) form hydrophobic and hydrophilic regions; they form a bilayer in water.
 - c. Steroids (a carbon skeleton with four fused rings) include cholesterol and sex hormones.
4. Proteins form diverse structures and functions.
 - a. A protein consists of one or more polypeptides, each a polymer of amino acids (from a set of 20).
 - b. A peptide bond joins the carboxyl group of one amino acid with the amino group of the next.

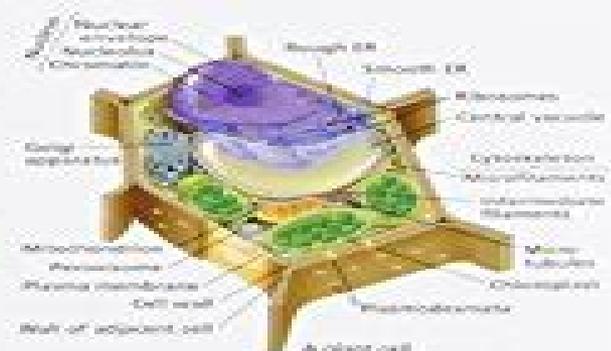
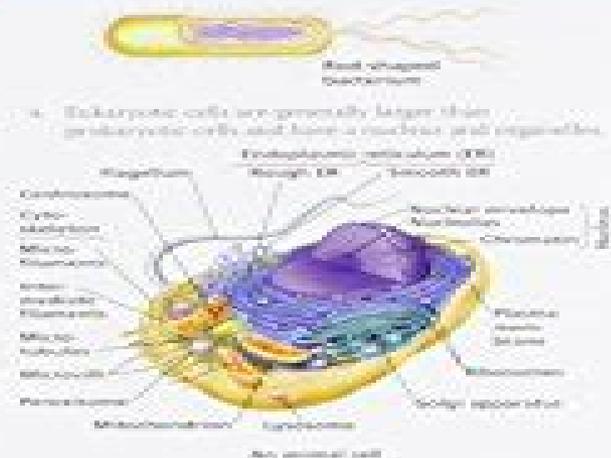
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- a. Four levels of protein structure lead to the unique shape and function of each protein.
2. Nucleic acids store, transmit, and help express hereditary information.
 - a. DNA and RNA are polymers of nucleotides.
 - b. A nucleotide has a nitrogenous base, a five-carbon sugar, and one or more phosphate groups.
 - c. DNA has deoxyribose as its sugar; RNA has ribose.
 - d. DNA has A, G, C, and T as bases; RNA has A, G, C, and U.
 - e. DNA is a double helix; RNA varies in shape.
 - f. A gene (hereditary unit) is a stretch of DNA that dictates the synthesis of RNA, which then directs the amino acid sequence of a polypeptide.

A Tour of the Cell (Ch. 6)

1. All organisms are composed of one or more cells.
2. All cells are bounded by a plasma membrane and contain cytosol, one or more chromosomes, and ribosomes.
3. Prokaryotic cells usually have a cell wall but lack a nucleus and membrane-bounded organelles.



3. Within the nucleus, chromosomes made up of a histone (DNA and proteins) carry the genetic instructions for making polypeptides.
4. Ribosomes are the cellular organelles that carry out protein synthesis.
5. The endomembrane system is a network of membranes, including the endoplasmic reticulum (ER).
 - a. Smooth ER has enzymes to lipid synthesis, carbohydrate metabolism, and detoxification.
 - b. Rough ER has ribosomes that synthesize secretory proteins, which enter the ER lumen. Rough ER also synthesizes membrane proteins and glycoproteins.
 - c. The Golgi apparatus refines ER products and packages them in vesicles.
 - d. Cytosomes, such as lysosomes, hydrolyze, enzymatically macromolecules and recycle organelles.
 - e. Vacuoles function in digestion, storage, waste disposal, water balance, and cell growth.
6. Mitochondria and chloroplasts change energy from one form to another.
 - a. Mitochondria, the sites of cellular respiration, extract energy from food and make ATP.

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Lingsheng Yao



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