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 $2 \text{ } 3 \text{ ca} + \text{oh} + 2 \text{ } 1 \text{ al} + 2 \text{ so}_4 \rightarrow 3 \text{ caso}_4 + 2 \text{ al oh} + 3$ type of reaction double displacement
 $3 \text{ } 3 \text{ mg} + 1 \text{ fe}_2\text{o}_3 \rightarrow 2 \text{ fe} + 3 \text{ mgo}$ type of reaction single displacement
 $4 \text{ } 1 \text{ c}_2\text{h}_4 + 3 \text{ o}_2 \rightarrow 2 \text{ co}_2 + 2$

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 $\text{naoh} + \text{kno}_3 \rightarrow \text{nano}_3 + \text{koh}$
 $\text{ch}_2\text{o} + \text{co}_2 \rightarrow \text{h}_4 + 2 \text{ } 2 \text{ } 2\text{o}_2 + \text{fe}_6 + \text{nabr}_2 + \text{febr}_6 + \text{na}_3 + \text{caso}_4 + \text{mg oh} + \text{ca oh} + \text{mgso}_4 + 2 \text{ } 2 \text{ } 4 \text{ nh}_4 + \text{oh} + \text{hbr} + \text{h}_2\text{o} + \text{nh}_4 + \text{br} + \text{pb} + \text{o}_2 + \text{pbo}_2 + \text{na}_2 + \text{co}_3 + \text{na}_2 + \text{o}_2 + \text{c}_4 + \text{h}_8 + \text{6o}_4 + \text{co}_2 + 2 \text{ } 4 \text{ h}_2 + \text{o} + \text{hcl} + \text{naoh} + \text{h}_2 + \text{nacl}$

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